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Formation Mechanism of Aluminum Nitride Powder via Organic Precursor Prepared from Anhydrous Aluminum Chloride and Ethylenediamine

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Synthesis of AIN powder via an organic precursor prepared from anhydrous aluminum chloride and ethylenediamine has been investigated to establish a simple and low-cost process. A brownish white solid product precipitated rapidly by simply mixing anhydrous aluminum chloride into ethylenediamine at room temperature in flowing dry N₂. Any crystallized compounds were not formed from the resulting organic precursor after calcination at 800°C for 5 h in flowing dry N₂, but formation of amorphous AIN was suggested from the FT-IR spectroscopic analysis. Crystallized AIN powder could be synthesized from the precursor after calcination at 1200°C for 5 h in flowing dry N₂. The crystallite size of AIN powder was less than 25 nm, but the particle size was in the range of 60-100 μm due to agglomeration of fine particles under the present synthesis conditions.

1. Introduction

Recent progress in technology for integrated circuits and power supply units has increased demand for silicon-compatible substrates with the thermal conductivity superior to that of conventional alumina substrates. Besides aluminum nitride (AIN), beryllium oxide and silicon carbide are possible candidates from the viewpoint of thermal conductivity. But, the use of these two materials is limited from potential toxicity in case of beryllium oxide and from high dielectric constant for silicon carbide relative to that of AIN. In addition to the high thermal conductivity, high electrical resistivity and thermal expansion coefficient comparable to that of silicon make AIN a more attractive material for the substrates, and therefore the application of AIN substrates has been expanding into various fields.

AIN powder is now prepared commercially either by direct nitridation of metallic aluminum or by carbothermal reduction-nitridation of alumina. However, grinding process is indispensable for the former, resulting in an increase in oxygen content in the powder. The increase in the oxygen content is the main cause for a decrease in thermal conductivity of AIN ceramics, and therefore new preparation process to reduce oxygen content in the resulting powder is highly requested. On the other hand, the latter requires high-purity alumina powder and reactions in the temperature range of 1500-2000°C, both of which lead to a rise in fabrication cost.

Thus, numerous efforts have so far been directed to developing new preparation methods of fine AIN powder. Typical examples are synthesis of fine AIN powder from Al powder by combined nitridation in fluidized-and packed-bed around 1400°C in pure N₂ and nitridation of Al powder in the microwave-induced N₂ plasma generated from SiC ceramics under atmospheric pressure. Various organic precursors have also been utilized to reduce temperatures for the synthesis of AIN powder. AIN powder was synthesized by the calcination of the polymeric precursor prepared from triethylaluminum and ethylenediamine (H₂NCH₂CH₂NH₂, ED) at 1000°C in an ammonia atmosphere. The use of hydrazine as a nitridation agent and lithium metal as a reductive agent

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for aluminum chloride reduced successfully the calcination temperature to 800°C for the synthesis of AIN powder even in an argon atmosphere. In addition, organic precursors prepared from the following binary and ternary systems have been studied for the synthesis of AIN powder: AlCl₃·6H₂O-(NH₂)₂CO, AlCl₃-glucose, LiAlH₄·PrNH₂, LiAlH₄·EtNH₂·HCl, and LiAlH₄·CH₃NH₂·HCl. Another interesting method is the formation an AIN polymer precursor, [Al(NH₂)(NH)₃], by electrolysis in a cell containing a NH₄Br/NH₃ electrolyte and Al electrodes at room temperature. AIN powder was then prepared by the calcination of the precursor above 600°C in flowing NH₃. However, most of these methods are characterized by their rather complicated procedure for the preparation of precursors. In addition, inflammable and expensive aluminum sources and ammonia as a nitrogen source are used in some cases, whereas simple and low-cost process is favorable from the viewpoint of industrial application.

The aim of the present study was to establish a simple and low-cost AIN synthesis process via an organic precursor prepared from anhydrous aluminum chloride and ethylenediamine in a N₂ atmosphere. A possible mechanism for the precursor formation and the AIN synthesis from the precursor has also been discussed based on the X-ray diffraction and fourier-transform infrared spectroscopic analyses of the products in each step of the synthesis.

2. Experimental

Anhydrous AlCl₃ (Wako Pure Chemical Industries, Ltd., reagent grade) and ED (Wako Pure Chemical Industries, Ltd., reagent grade) were used as an aluminum source and a nitrogen source, respectively, without further purification. The liquid ED was placed in a three-neck flask, and the atmosphere inside the flask was purged with flowing dry N₂ (99.99%). Anhydrous AlCl₃ was then added slowly to the liquid ED up to a molar ratio of AlCl₃ : ED = 1 : 10 at room temperature while stirring and flowing dry N₂. The resulting suspension smoked along with a rise in its temperature during the addition, and then a brownish white solid product was precipitated rapidly. The solid product was dried at 150°C for about 2 h in flowing dry N₂ to remove excess ED, subsequently treated in vacuo at 100°C for 3 h.

The precursor powder thus obtained was calcined at 800 or 1200°C for 5 h in a tubular-type electric furnace in flowing dry N₂. Carbon fibers were placed at the upper stream in a reaction tube and was used to eliminate trace O₂ in flowing dry N₂ during the calcination.

The powders before and after the calcination were characterized by X-ray diffraction (XRD, Rigaku, RINT-2200), fourier-transform infrared spectroscopy (FT-IR, Perkin Elmer, System 2000) and scanning electron microscopy (SEM, Hitachi, S-2250N).

3. Results and Discussion

3.1 Characterization of As-dried Organic Precursor

Figure 1 shows an XRD pattern of the as-dried organic precursor synthesized from AlCl₃ and ED. The XRD pattern of the precursor was completely different from that of AlCl₃, and could not be identified with the available JCPDS (Joint Committee on Powder Diffraction Standards) cards. This implied the formation of a new compound by the reaction of AlCl₃ with ED. Therefore, FT-IR spectroscopic analysis was carried out to get information on the chemical bonds involved in the precursor based on the data in Ref. 13. Since the peaks around 3500 and 3400 cm⁻¹ generally stand for

![Fig. 1 XRD pattern of the as-dried organic precursor synthesized from AlCl₃ and ED.](image)

![Fig. 2 FT-IR spectrum of the as-dried organic precursor synthesized from AlCl₃ and ED.](image)
3.2 Characterization of AIN Powder after Calcination of Precursor

Figures 4 and 5 show XRD and FT-IR spectra of the resulting powders obtained after calcination at 800 and 1200°C for 5 h in flowing dry N₂, respectively. The powder after calcination at 800°C was black in color and showed a featureless XRD pattern as shown in Fig. 4(a). But, a small peak ascribed to Al-N bonds are observed in Fig. 5(a). In addition, absorption peaks around 3500 and 3150 cm⁻¹ almost disappeared, and those in the ranges of 1300-1600 (corresponding to N-H bending), 1000-1300 (C-N stretching) and 500-1000 cm⁻¹ (N-H out-of-plane vibrations) weakened significantly (compare Fig. 5(a) with Fig. 2). These results suggest decomposition and evaporation of organic groups in the precursor during the calcination at 800°C in N₂, and the resulting powder seems to be a mixture of amorphous AIN and carbonaceous compounds.

In contrast, formation of a hexagonal and single AIN phase was confirmed after the calcination at 1200°C, as shown in Fig. 4(b-1), whereas the resulting powder was still black in color. The color of the powder turned into gray after subsequent firing at 600°C for 2 h in flowing air, but no oxide phase was induced by the firing, as shown in Fig. 4(b-2). The FT-IR spectrum of the powder after the firing shown in Fig. 5(b) exhibited a strong adsorption peak around 700 cm⁻¹ and quite resembled that of AIN powder supplied from Tokuyama Corp. (F-grade) in Fig. 5(c). In addition, absorption peaks assigned to N-H bending, C-N stretching or N-H out-of-plane vibrations were hardly observed after the firing. Thus, it is confirmed that the AIN powder after the calcination at 1200°C in N₂ still contains some carbonaceous compounds, but these compounds can be removed by the subsequent firing.

The crystallite size of AIN powder, calculated by Scherrer’s equation, was 14.6 nm after the calcination at 1200°C in N₂ and remained unchanged after the subsequent firing. Owing to such fine microstructure of the synthesized AIN powder, its surface is likely hydrolyzed during handing. A broad peak at 3200-3600 cm⁻¹ and another around 1650 cm⁻¹ observed for the powder after the firing (see Fig. 5(b)) are, therefore, probably due to adsorbed water.

Deposition of a paste-like yellow material was observed at the outlet of the reaction tube after the
Calcined at 1200°C in N₂ followed by firing at 600°C in air.

Crystallite size: 14.0 nm

Fig. 6 XRD pattern of the material deposited at the outlet of the reaction tube after the calcination of precursor at 1200°C for 5 h in N₂.

(a) Calcined at 800°C in N₂

(b) Calcined at 1200°C in N₂

Crystallite size: 14.6 nm

Fig. 4 XRD patterns of resulting powders after calcination of precursor.

Fig. 7 Possible chemical reactions for the synthesis of AlN powder.

Calcination of the precursor at both 800 and 1200°C in N₂. XRD analysis of the compound revealed a mixture of H₂NCH₂CH₂NH₂·2HCl and NH₄Cl, as shown in Fig. 6. Then, it is confirmed that the thermal decomposition of the precursor proceeds by evolving ED and HCl. Based on the XRD and FT-IR analyses, possible chemical reactions taking place during the synthesis of AlN powder from the precursor are schematically shown in Fig. 7. The precursor is confirmed to have some Al-Cl bonds in the polymerized structure, as shown in Fig. 7(a). Polymerization is further accelerated upon heating, being accompanied with the evolution of HCl, as shown
in Fig. 7(b). Network formation shown in Fig. 7(c) is completed by the evolution of ED. The powder after the calcination at 800°C seems to have such network structure as shown in Fig. 7(c), since it exhibited absorption peaks assigned to N-H bending, C-N stretching or N-H out-of plane vibrations in the FT-IR spectrum. Heating at higher temperature, i.e. 1200°C in the present study, leads to the release and/or the decomposition of R-NH groups in the network structure, and results in the formation of AlN crystallites. \( \text{H}_2\text{NCH}_2\text{CH}_2\text{NH}_2 \cdot 2\text{HCl} \) is undoubtedly formed by the reaction between the evolved ED and HCl, and NH4Cl may be formed by the decomposition of R-NH groups and HCl. Some carbonaceous compounds are produced by the heating of R-NH groups in N2 and are the origin of the black-colored products after the calcination.

Figure 8 shows SEM photographs of the AlN powder prepared from the precursor by the calcination at 1200 °C for 5 h in N2 and the subsequent firing at 600°C for 2 h in air. It is confirmed from the low and high resolution photographs that the powder consists of very fine primary particles less than 25 nm, and this particle size is almost comparable to the crystallite size calculated from the XRD data. However, it is obvious that these fine particles are agglomerated to form a secondary particle in the range of 60-100 μm. Such feature may reduce the sinterability of the AlN powder. In this study, no grinding process of the precursor was adopted before the calcination, which is of course one reason for the formation of such large secondary particles. However, the rapid solidification of the precursor in ED solution is anticipated to be the main reason. To find milder reaction conditions for the preparation of the precursor, e.g. by employing a suitable solvent, and then to obtain finer secondary AlN particles are our next goal in future work.

4. Conclusions

The organic precursor capable of producing AlN powder has been prepared by simply mixing anhydrous aluminum chloride into ethylenediamine at room temperature in flowing dry N2. Calcination of the precursor at 1200 °C for 5 in flowing dry N2 led to a mixture of crystallized AlN powder and carbonaceous compounds, whereas amorphous AlN was formed after calcination at 800°C for 5 h. The carbonaceous compounds could be removed by subsequent firing at 600°C for 2 h in air, without any changes in crystal structure and crystallite size of the AlN powder. Possible chemical reactions during the formation of the precursor and AlN crystallites have been proposed based on the XRD and FT-IR spectroscopic analyses of both the products and by-products. The crystallite size of the AlN powder was less than 25 nm, but the particle size was in the range of 60-100 μm due to agglomeration of fine particles under the present synthesis conditions.

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